



**CALIFORNIA STATE SCIENCE FAIR  
2016 PROJECT SUMMARY**

<b>Name(s)</b> Berenice Vega	<b>Project Number</b> <b>J0634</b>
<b>Project Title</b> <b>Reaction Rates: Does Temperature Affect the Iodine Clock Reaction?</b>	
<b>Abstract</b> <b>Objectives/Goals</b> The goal was to find out if temperature affected the rate of the iodine clock reaction. <b>Methods/Materials</b> Stopwatch, distilled water, Vitamin C tablets, 2% Iodine tincture, hydrogen peroxide, liquid laundry starch, disposable cups, thermometer, gas stove, an refrigerator. I made two solutions and poured them at the same time while timing the reaction. Three different batches were made one hot, one cold, and one room temperature. <b>Results</b> After many trials to ensure accuracy, it was evident that temperature played a key role in reaction rates. The solutions made with cold water were the slowest to react, while the solutions made with hot water were the fastest to react. <b>Conclusions/Discussion</b> After determining the results of the experiment, it is evident that temperature plays a key role in reaction rates. This supports the idea that you can alter the amount of time it will take for a chemical reaction to react by varying the temperature.	
<b>Summary Statement</b> Based on the results, temperature plays a key role in altering the reaction time of the Iodine Clock reaction.	
<b>Help Received</b> I performed the Iodine Clock reaction myself, however, I received aid in understanding what happens in the reaction from the website Science Bob.	